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## Chemical Equilibrium

Equilibrium! Chapter 15 – Chemical  
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Section

CHAPTER 18 REVIEW Chemical

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## Chemical Equilibrium

Equilibrium CHAPTER 18 REVIEW

Chemical Equilibrium SECTION 1

SHORT ANSWER Answer the

following questions in the space

provided. 1. Write the equilibrium

expression for the following

hypothetical equation:  $3A(aq) + B(aq)$

$2C(aq) + 3D(aq)$  2. a. Write the

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appropriate chemical equilibrium  
expression for each of the

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Chapter 18 Review Chemical  
Equilibrium Answers

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Equilibrium CHAPTER 18 REVIEW

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## Chemical Equilibrium

### Chemical Equilibrium SECTION 1

SHORT ANSWER Answer the following questions in the space provided.

1. Write the equilibrium expression for the following hypothetical equation:  $3A(aq) + B(aq) \rightleftharpoons 2C(aq) + 3D(aq)$

2. a. Write the appropriate chemical equilibrium

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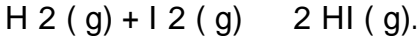
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## Chemical Equilibrium

### Equilibrium Mixed Answers ...

Figure 9.5.1: Equilibrium in reaction:



Chemical equilibrium can be attained whether the reaction begins with all reactants and no products, all products and no reactants, or some of both. The figure below shows changes



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## Chemical Equilibrium

Answers to questions about the concentration of  $H_2$ ,  $I_2$ , and  $HI$  for two different reactions.

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## Chemical Equilibrium

Equilibrium CHAPTER 18 REVIEW

Chemical Equilibrium SECTION 2

SHORT ANSWER Answer the following questions in the space provided. 1. \_\_\_\_\_ Raising the temperature of any equilibrium system always (a) favors the forward reaction. (b) favors the reverse

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Answers Section 2  
reaction. (c) favors the exothermic reaction.

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Review Chemical Equilibrium Section  
2

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a state of balance in which the rates of  
the forward and reverse reactions are

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## Chemical Equilibrium

Answers Section 2  
equal; no net change in the amount of reactants and products occurs in the chemical system (18.2) equilibrium position the relative concentrations of reactants and products of a reaction that has reached equilibrium; indicates whether the reactants or products are favored in the reversible



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Section 2

Equilibrium Section Review

Answersthe forward and reverse  
reactions are equal; no net change in  
the amount of reactants and products

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## Chemical Equilibrium

Answers Section 2  
occurs in the chemical system (18.2)  
equilibrium position the relative  
concentrations of reactants and  
products of a reaction that has  
reached equilibrium; indicates  
whether the reactants or products are

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SECTION. .4? SHORT

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## Chemical Equilibrium

Chapter 18 Section 4 Review - Name  
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Consider the following equation:

$2\text{NO}(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{NO}_2(\text{g})$  At  
equilibrium,  $[\text{NO}] = 0.80 \text{ M}$ ,  $[\text{O}_2] =$   
 $0.50 \text{ M}$ , and  $[\text{NO}_2] = 0.60 \text{ M}$ . Calculate  
the value of  $K$ . 4. Consider the  
following equilibrium equation:



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## Chemical Equilibrium



a. Increasing  $[\text{CO}]$  will (a) increase  $[\text{H}_2]$ . decrease  $[\text{H}_2]$ . (c) not change  $[\text{H}_2]$ .

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## Chemical Equilibrium

View Answer. Section 2  
The equilibrium system below shows the reaction of sour gas with methane to produce carbon disulfide and hydrogen gas.  $\text{CH}_4 (\text{g}) + 2\text{H}_2\text{S} (\text{g}) \rightleftharpoons \text{CS}_2 (\text{g}) + 4\text{H}_2 (\text{g})$ ;  $\Delta H = +232.5 \text{ kJ}$ ...

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## Chemical Equilibrium

Answers Section 2  
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Solution: In systems 1 and 3, the concentration of A decreases from  $t_0$  through  $t_2$  but is the same at both  $t_2$  and  $t_3$ . Thus systems 1 and 3 are at equilibrium by  $t_3$ . In system 2, the concentrations of A and B are still

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## Chemical Equilibrium

changing between  $t_2$  and  $t_3$ , so system 2 may not yet have reached equilibrium by  $t_3$ .

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5.1: The Concept of Equilibrium -  
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chemical equilibrium because its

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## Chemical Equilibrium

Chemical equation is written with a double arrow like this.  $N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g)$  At equilibrium, the concentrations of reactants and products are constant, as you saw in Figures 18-2c and 18-2d. However, that doesn't mean that the amounts or concentrations of reactants and

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## Chemical Equilibrium

products are equal. That is sel-dom  
the case.

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